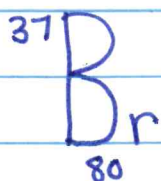
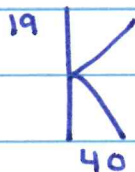


## Assessment Card



P-37  
N-45  
E-37

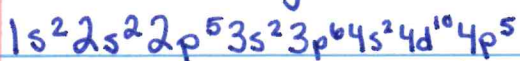


P-19  
N-22  
E-20

- has the largest atomic radius

- has a high electronegativity

- Electron Configuration:



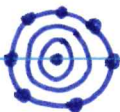
- Lewis Dot Structure:  $\cdot\ddot{\text{Br}}\cdot$

- Bromine is a noble gas

- Bromine has 17 valence  $e^-$

- Bromine has a low ionization energy

- Bohr model:



- has the smallest atomic radius

- has the highest electronegativity

- Noble Gas Configuration:



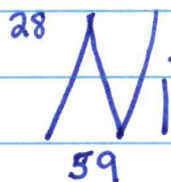
- Lewis Dot structure:  $\ddot{\text{K}}$

- Potassium reacts violently with  $\text{H}_2\text{O}$

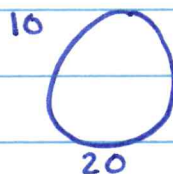
- Potassium has one valence  $e^-$

- Potassium is an alkaline earth metal

- Bohr model:



P-28  
N-31  
E-28

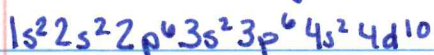


P-8  
N-8  
E-8

- has a high ionization energy

- has <sup>the</sup> smallest atomic radius

- Electron Configuration:



- Nickel is an alkali metal

- Arrow notation:



- has a high ionization energy

- has a small atomic radius

- Arrow configuration:  $\uparrow\downarrow \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$

- Lewis Dot structure:  $\ddot{\text{O}}\cdot$

- Oxygen has 7 valence  $e^-$

- Oxygen electron configuration:



- Bohr Model:

