

watch the videos on my website under unit 6: VSEPR Theory

## VSEPR Theory

Define VSEPR-

Key-

A- Central atom

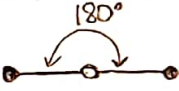
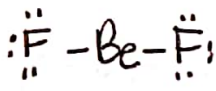

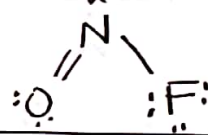

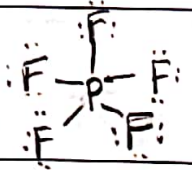
B- Bonded atom

E- Lone pair of electrons

○ -

● -

⦿ - Lone pair of electrons.

Name of Shape	Actual Shape	# of Atoms Bonded to the Central Atom	# of Lone Pairs of Electrons on the Central	Basic Molecular Formula	Actual Formula Example	Lewis Structure
1 <i>example</i> Linear		2	0	AB <sub>2</sub>	BeF <sub>2</sub>	
2		3	0			
3		2	1	AB <sub>2</sub> E	ONF	
4					CH <sub>4</sub>	
5				AB <sub>3</sub> E		
6 Bent			2			
7						
8	all 90° angles	6	0			